

Name: _____ Class: _____ Date: _____

Assessment

Chemical Formulas and Chemical Compounds

Teacher Notes and Answers

SECTION: OXIDATION NUMBERS

1. c
2. b
3. a
4. c
5. b
6. d
7. d
8. b
9. c
10. a

Assessment

Chemical Formulas and Chemical Compounds

Section Quiz: Oxidation Numbers

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. An oxidation number
- is always negative.
 - is always positive.
 - has no exact physical meaning.
 - is the same in each compound that an element forms.
- _____ 2. Which element has the same oxidation number in all of its compounds?
- oxygen
 - fluorine
 - hydrogen
 - chlorine
- _____ 3. What is the oxidation number of an atom of nitrogen in the compound N_2 ?
- 0
 - +3
 - 3
 - 6
- _____ 4. What is the oxidation number of phosphorus in H_3PO_4 ?
- +1
 - +4
 - +5
 - +8
- _____ 5. What is the oxidation number of As in AsO_4^{3-} ?
- +8
 - +5
 - 5
 - 8

Section Quiz, *continued*

- _____ 6. What is the correct Stock name for $\text{Cr}(\text{CH}_3\text{COO})_3$
- chromium acetate
 - chromium triacetate
 - trichromium acetate
 - chromium(III) acetate
- _____ 7. What is the correct prefix-based name for selenium(II) fluoride?
- selenium fluoride
 - diselenium fluoride
 - diselenium difluoride
 - selenium difluoride
- _____ 8. What is the formula for copper(I) oxide?
- CuO
 - Cu_2O
 - CuO_2
 - Cu_2O_2
- _____ 9. A monatomic ion has an oxidation number
- of zero.
 - indicated by the ion's prefix.
 - equal to the charge of the ion.
 - equal to its subscript.
- _____ 10. What is the correct Stock name for SnBr_4 ?
- tin(IV) bromide
 - tin bromide
 - tin tetrabromide
 - tin(IV) tetrabromide.